INTRODUCTION TO BIOCHEMISTRY

A STUDENT SHOULD BE ABLE TO:

1. With respect to lipids, know:
   The characteristic common to members of the class: solubility in nonpolar solvents
   The functional groups most important in triglycerides (triacylglycerols): esters, and in
   unsaturated fats, alkenes
   The characteristic common to most naturally occurring fatty acids: even numbers of
   carbon atoms; \textit{cis} configuration for unsaturated compounds; lack of branching
   Saturated, unsaturated, and polyunsaturated fats; \textit{cis} and \textit{trans} fatty acids
   Characteristics that lead to higher melting points for fatty acids: saturation and higher
   molecular weight; \textit{trans over cis}
   The products (soap and glycerol) of saponification reactions given the triglyceride
   starting material, or \textit{vice versa}
   Structural features common to terpenes, steroids, and prostaglandins

2. With respect to carbohydrates, know:
   The approximate empirical formula characteristic of members of the class: CH\textsubscript{2}O
   The functional groups present in carbohydrate molecules: alcohols (usually several
   are present); aldehydes and ketones; hemiacetals and acetics
   Carbohydrates belonging to these classes: ketoses and aldoses; D and L sugars;
   monosaccharide classes (trioses, tetroses, pentoses, hexoses, etc.); \textit{α} and \textit{β}
   pyranoses; reducing and nonreducing sugars
   The products of oxidation and reduction reactions of carbohydrates (reducing sugars
   are those that can be oxidized. They contain hemiacetals (in the cyclic form) and
   aldehydes or ketones (in the straight chain form)
   The number of chirality centers in a carbohydrate, given a Fisher projection
   How to draw the cyclic form given the acyclic structure, and acyclic given the cyclic

3. With respect to amino acids, polypeptides, and proteins, know:
   The functional groups most important in these compounds. Proteins and polypeptides
   contain amides made from amino acid monomers; there are carboxylic acid and
   amine groups on the ends.
   The degree and site of protonation of amino acids as a function of pH
   Relative isoelectric points, pI
   Polypeptide structures from N-terminal analysis, C-terminal analysis, and total and
   partial hydrolysis results
   How to link amino acids together to make polypeptides
   Synthesis of amino acids \textit{via} \textit{α}-haloacids and Strecker synthesis
To best prepare for this module, please work suitable Chapters 24-26 Skill Builder problems in the textbook.

A STUDENT WHO HAS MASTERED THE OBJECTIVES FOR THIS UNIT SHOULD BE ABLE TO SOLVE THE FOLLOWING PROBLEMS AND RELATED ONES:

1.1 In each of these pairs, which of the compounds has the higher melting point?
   a) CH₃(CH₂)₁₆COOH or CH₃(CH₂)₇CH=CH(CH₂)₇COOH
   b) CH₃(CH₂)₁₄COOH or CH₃(CH₂)₁₆COOH

1.2 Which of the following fatty acids is a) saturated, b) monounsaturated, and c) polyunsaturated?
   a) CH₃(CH₂)₁₄COOH
   b) CH₃(CH₂)₄(CH=CHCH₂)₄(CH₂)₂COOH
   c) CH₃(CH₂)₇CH=CH(CH₂)₇COOH

1.3 a) Which of the following fatty acids is found commonly in cells?
   a)
   b)
   c)
   d)

   b) Identify each of the compounds in part (a) as cis, trans, or neither.

1.4 Match the structures shown below to the appropriate type of lipid:

   a) prostaglandin
   b) triglyceride
   c) terpene
   d) steroid
2.1 For each of the structures below, provide the information needed to complete the table.

<table>
<thead>
<tr>
<th>Structure</th>
<th>Family</th>
<th>Backbone</th>
<th>Configuration</th>
<th># of Chirality Centers</th>
</tr>
</thead>
<tbody>
<tr>
<td>a)</td>
<td>ketose</td>
<td>tetrose</td>
<td>D</td>
<td>1</td>
</tr>
<tr>
<td>b)</td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>c)</td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>d)</td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>e)</td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
</tbody>
</table>

2.2 a) Which of these disaccharides, shown here in different drawing conventions, gives a negative Tollens’ test (that is, does not give a silver mirror when treated with a solution of silver nitrate in aqueous ammonia)?

b) Why would heating the compound that is the answer to part (a) with dilute aqueous acid yield a product that gives a positive Tollens’ test?

![Structures](image-url)
2.3 Draw the β-pyranose form of talose:

![β-pyranose form of talose diagram]

2.4 The structure of an important plant ketopentose, ribulose, is shown below. How many chirality centers does this molecule contain? Draw the structures of the other stereoisomers, and identify each as D or L.

![Structure of ribulose]

3.1 Based on the pKₐ values given, draw the dominant form of each of the following amino acids at pH 3, at pH 7, and at pH 13. Also, calculate the pI for each.

a) Glycine: pKₐ values of 2.3 (for the α-COOH) and 9.6 (for the α-NH₃⁺).

![Glycine structure]

b) Glutamic acid, pKₐ values of 2.2, 9.7, and 4.3 (for the COOH group at the bottom).

![Glutamic acid structure]
3.1  c) Histidine, pK\textsubscript{a} values of 1.8, 9.2, and 6.0 (for the conjugate acid of the N designated with a *).

<table>
<thead>
<tr>
<th>pH 3</th>
<th>pH 7</th>
<th>pH 13</th>
<th>pI</th>
</tr>
</thead>
</table>

\[
\text{CO}_2\text{H} \quad \text{NH}_2 \quad \text{H} \quad \text{CH}_2 \quad \text{N} = \text{NH} \\
\text{H} \quad \text{H} \quad \text{H} \quad \text{H} \quad \text{H}
\]

3.2. A pentapeptide was completely hydrolyzed and found to contain aspartic acid (Asp), lysine (Lys), proline (Pro), and methionine (Met). After partial hydrolysis the dipeptides Lys-Asp, Met-Pro, Lys-Met, and Pro-Lys were identified in the product. What is the sequence of amino acids in this pentapeptide?

a) Lys-Met-Pro-Lys-Asp  
b) Asp-Lys-Pro-Met-Lys  
c) Lys-Asp-Met-Pro-Lys  
d) Pro-Met-Lys-Pro-Asp

3.3 A tridecapeptide (an oligopeptide with 13 amino acid residues) has the following amino acid composition: Ala, Arg, Asp(2), Glu(2), Gly(3), Leu, and Val(3). After partial acid hydrolysis of the tridecapeptide, the following peptides were isolated, and their sequences were determined by Erdman degradation. What is the sequence of the original tridecapeptide?

The peptides are:
- Asp-Glu-Val-Gly-Gly-Glu-Ala
- Val-Asp-Val-Asp-Glu
- Val-Asp-Val
- Glu-Ala-Leu-Gly-Arg
- Val-Gly-Gly-Glu-Ala-Leu-Gly-Arg
- Leu-Gly-Arg

3.4 Draw the structure that results when Asp, Gly, Glu and Ala are joined together to form this peptide: Asp-Gly-Glu-Ala. The structures of the amino acids are in your book.
3.5 Predict the product(s) of these reactions:

a) \[
\text{Br}_2, \text{PBr}_3 \rightarrow \text{Br}_3 \rightarrow \text{xs NH}_3
\]

b) \[
\text{NH}_4\text{Cl, NaCN} \rightarrow \text{H}_3\text{O}^+
\]

3.6 Synthesize alanine (CH$_3$CH(NH$_2$)CO$_2$H) starting from:

a) propanoic acid

b) acetaldehyde

SOLUTIONS TO SAMPLE PROBLEMS:

1.1 a) CH$_3$(CH$_2$)$_{16}$COOH  
    b) CH$_3$(CH$_2$)$_{16}$COOH

1.2 a) a  
    b) c  
    c) b

1.3 a) b  
    b) b and d are cis, c is trans, and a is neither cis nor trans.

1.4 a) steroid (cortisone)  
    b) prostaglandin  
    c) terpene (camphor)

2.1 Structure Family Backbone Configuration Number
a) ketose tetrose D 1
b) ketose hexose D 3
c) aldose hexose L 4
d) ketose triose none 0
e) aldose pentose D 3

2.2 a) Sucrose would give a negative Tollens’ test. This indicates that sucrose is not a reducing sugar because sucrose does not contain a hemiacetal. Lactose would give a positive Tollens’ test due to the hemiacetal functional group contained in the structure. The cyclic hemiacetal ring can open to the carbonyl, which undergoes oxidation.

b) Warming with dilute aqueous acid hydrolyzes sucrose into two monosaccharides by cleaving the acetal linkage. These monosaccharides could then cyclize to produce hemiacetal groups, and thus they become reducing sugars.
2.3  

![Chemical structure diagram]

Groups on the left of the Fisher projection are up in the cyclic form.

2.4  

Ribulose contains two chirality centers. The structures of the isomers are:

- **D** ribulose
- **L** ribulose

3.1  

a) Glycine pKₐ values: 2.4 and 9.8. If the pKₐ value for a given proton is less than the pH, that proton is removed. pI = (2.4 + 9.8)/2 = 6.1

- **Glycine**
  - pH = 3
  - pH = 7
  - pH = 13

b) Glutamic acid: pI = average of pKa values for two acidic groups = 3.25

- **Glutamine**
  - pH = 3
  - pH = 7
  - pH = 13
3.1 c) Histidine: pI = average of pKa values for two basic groups = 7.6

\[
\begin{align*}
\text{Histidine} & : \quad \text{pH} = 3, 7, 13 \\
\text{pK}_{\alpha} & = 2.2, 7.6, 10.5
\end{align*}
\]

3.2 a

3.3 Val-Asp-Val-Asp-Glu-Val-Gly-Gly-Glu-Ala-Leu-Gly-Arg

3.4

\[
\begin{align*}
\text{N-terminus} & : \quad \text{Asp} \quad \text{Gly} \\
\text{C-terminus} & : \quad \text{Glu} \quad \text{Ala}
\end{align*}
\]

3.5 Predict the product(s) of these reactions:

\[
\begin{align*}
a) & : \quad \text{CH}_3\text{CH(OH)} \quad 1. \text{Br}_2, \text{PBr}_3 \rightarrow \text{CH}_3\text{CHBr} \quad 2. \text{H}_2\text{O} \rightarrow \text{CH}_3\text{CHNH}_3^+ \\
& \text{b) } \text{NH}_4\text{Cl} \quad \text{NaCN} \rightarrow \text{H}_2\text{NCH}_2\text{CN} \quad \text{H}_3\text{O}^+ \rightarrow \text{H}_3\text{NCH}_2\text{CO}_2\text{H}
\end{align*}
\]

3.6 Synthesize alanine (CH\(_3\)CH(NH\(_2\))CO\(_2\)H) starting from:

\[
\begin{align*}
a) & : \quad \text{CH}_3\text{COOH} \quad 1. \text{Br}_2, \text{PBr}_3 \rightarrow \text{CH}_3\text{CHBr} \quad 2. \text{H}_2\text{O} \rightarrow \text{CH}_3\text{CHNH}_3^+ \\
& \text{b) } \text{NH}_4\text{Cl} \quad \text{NaCN} \rightarrow \text{H}_2\text{NCH}_2\text{CN} \quad \text{H}_3\text{O}^+ \rightarrow \text{H}_3\text{NCH}_2\text{CO}_2\text{H}
\end{align*}
\]
1. What is the structure of the amino acid HO$_2$CCH$_2$CH(NH$_2$)COOH at a pH of 2.5? The pK$_a$ values are 2.2, 9.7, and 4.3 (for the second (bottom) COOH).

   a) \[ \text{\text{\begin{tabular}{c}
   \text{COO} \\
   \text{H} \\
   \text{\text{H}$_3$N} \\
   \text{\text{H}$_2$N} \\
   \text{CH$_2$CO$_2$} \\
   \end{tabular}}} \]
   b) \[ \text{\text{\begin{tabular}{c}
   \text{COO} \\
   \text{H} \\
   \text{\text{H}$_3$N} \\
   \text{\text{H}$_2$N} \\
   \text{CH$_2$CO$_2$H} \\
   \end{tabular}}} \]
   c) \[ \text{\text{\begin{tabular}{c}
   \text{COOH} \\
   \text{H} \\
   \text{\text{H}$_3$N} \\
   \text{\text{H}$_2$N} \\
   \text{CH$_2$CO$_2$H} \\
   \end{tabular}}} \]
   d) \[ \text{\text{\begin{tabular}{c}
   \text{COO} \\
   \text{H} \\
   \text{\text{H}$_3$N} \\
   \text{\text{H}$_2$N} \\
   \text{CH$_2$CO$_2$H} \\
   \end{tabular}}} \]

2. Which of the following functional groups is NOT present in the molecule shown?

   a) Acetal   b) Hemiacetal   c) 1° Alcohol   d) 2° Alcohol

3. What functional group is most abundant in proteins?

   a) Amide   b) Amine   c) Carboxylic acid   d) Alcohol

4. What are the major products of the reaction shown?

   \[
   \text{CH}_3(\text{CH}_2)_{12}\text{CO}_2\text{CH} \xrightarrow{\text{NaOH excess}} \text{CH}_3(\text{CH}_2)_{12}\text{CO}_2\text{CH}_2
   \]

   I. CH$_2$—CH—CH$_2$   II. CH$_2$—CH—CH$_2$   III. CH$_3$(CH$_2$)$_{12}$CO$_2$H   IV. CH$_3$(CH$_2$)$_{12}$CO$_2$Na

   a) I and III   b) I and IV   c) II and III   d) II and IV

5. The structure of ribose is shown. Ribose is a (an):

   a) Aldopentose   b) Aldohexose   c) Ketopentose   d) Ketohexose
1. Complete the following equation.

\[ \text{HO} - \text{O} \quad \text{NH}_2\text{Cl} \quad \text{NaCN} \quad \text{H}_2\text{O}^+ \]

2. Which of these compounds does NOT give a red precipitate with Benedict’s reagent?

   a) ![Compound A]
   b) ![Compound B]
   c) ![Compound C]
   d) ![Compound D]

3. What is the structure of the amino acid shown at a pH of 2? The pKa values are 2.2 (COOH), 9.0 (NH₂), and 10.5 (side chain NH₂). What is the pI?

   ![Amino Acid Structure]

4. An oligopeptide was analyzed and found to contain the amino acids Phe, Cys, Ala, Ile, Asp, Trp, Thr, and Arg. N-terminal analysis identified Phe, and C-terminal analysis gave Ile. The fragments Arg-Ala-Ile, Cys-Asp-Arg, and Trp-Thr-Cys were identified after partial hydrolysis. What is the amino acid sequence of this peptide?
Learn the basics of biochemistry with this introduction and test your knowledge with exam questions! Learn more about how living organisms arose from biomolecules. Diversity and spread of life ✓, historical & technological improvement ✓, the avery-macleod-mccarty experiment ✓. Start now! Lecturio is using cookies to improve your user experience. By continuing use of our service you agree upon our Data Privacy Statement.

Introduction to Biochemistry

What Does a Biochemist Do? Many biochemists work in chemistry labs. Helmenstine, Anne Marie, Ph.D. “Biochemistry Introduction and Overview.” ThoughtCo, Aug. 26, 2020, thoughtco.com/biochemistry-introduction-603879. Helmenstine, Anne Marie, Ph.D. (2020, August 26). 1 Introduction to Biochemistry And Mol & Biol Prof. Shaoping Ji Department of Biochemistry and Molecular Biology Medical School Henan University. 2 What is Biochemistry? Biochemistry is the application of chemistry to the study of biological processes at the cellular and molecular level. Introduction to Biochemistry. Biochemists discuss chemistry with biologists, and biology with chemists, thereby confusing both groups. Among themselves, they talk about baseball. Credit:Anonymous. As the name indicates, biochemistry is a hybrid science: Biology is the science of living organisms and chemistry is the science of atoms and molecules, so biochemistry is the science of the atoms and molecules in living organisms.